

task_4pla882qt21jfyc9_with_calculation

Student Group

First Name	Surname	Matrikel Nr.

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Exercise E1 Charges in Electroplating

To get a different metal coating onto a surface, often  **Electroplating** is used. In this process, the surface is located in a liquid, which contains metal ions of the coating.

In the following, a copper coating (e.g. for corrosion resistance) shall be looked on. The charge of one copper ion is around $1.6022 \cdot 10^{-19} \text{~}\mathrm{C}$, what is the charge on the surface if there are $8 \cdot 10^{22} \text{~}\mathrm{ions}$ added?

$$12'818 \text{~}\mathrm{C}$$

Solution

$$8 \cdot 10^{22} \cdot 1.6022 \cdot 10^{-19} \text{~}\mathrm{C} = 12'817.6 \text{~}\mathrm{C}$$

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